Chemistry 120 Week 2 Quiz 1 questions

1) You have a sample with a mass of 25 g and a volume of 2.8 ml. What is the density of this sample?

D = M/V m= 25g v= 2.8 mL X 0.001 dm3 / 1 mL V= 2.801 dm3

v= 2.8mL D= 25g / 2.801 dm3

Answer: 8.9 g/mL

2) How name valence electrons are there in an Aton of Neon?

you look at to column at the periodic table. Ne is on column 8

Answer: 8

3) How name pounds (lbs) are in 58kg?

Convert: 2.2 lbs = 1 kg you'll multiply 2.2 X 58 = 127.6

Answer: 127.6 lbs

3) Which atom would have an electron configuration of 1s2 2s2 2p6 3s2

3p6? Look at the periodic table and count

Answer: Ar

4) Determine how many electrons (e-) you would expect of the following ions and ground state atoms to contain and match them up.

+ charge = Cations (lost e) - charge = Anions (gain e)

b) Fluoride - 10

c) Flourine - 9

d) Zn (I) - 29

e) Calcium ion- 18

5) A ground state atom (uncharged atom) gains 3 electrons. This atom will now have a charge of:

Answer: -3

6) What is the correct Lewis Dot Structure for Phosphorous?